

## ***Interactive comment on “Application of a laser-based spectrometer for continuous insitu measurements of stable isotopes of soil CO<sub>2</sub> in calcareous and acidic soils” by Jobin Joseph et al.***

**Jobin Joseph et al.**

jobinjoseph224@gmail.com

Received and published: 30 September 2018

### **Reviewer 1# Remarks to the Author**

This manuscript describes attempts to calibrate a laser-based absorption instrument for use in high-frequency measurements of <sup>13</sup>C and <sup>18</sup>O of CO<sub>2</sub> in soil depth profiles, and provides very brief field data from two sites. This type of work is useful in that many investigators use new instruments, such as the LGR instrument described here, without sufficient validation. However, I have some major concerns about the calibration method which the authors can hopefully address in a revised manuscript. In addition, the paper would be much stronger if additional field data were presented, especially

C1

along with atmospheric measurements at the soil surface, which are needed to calculate the isotope composition of soil respired CO<sub>2</sub> (as opposed to soil profile CO<sub>2</sub>). At present, it appears that <24 hours of field data are shown.

1. The authors develop non-linear calibration functions to account for the concentration dependence of isotope ratios but it is not clear to me how these functions might also vary in isotope space (i.e., as a dual function of isotope composition and concentration). For example, fig. 5 shows the correction functions for concentration dependence but does not show how/if these varied as a function of the isotope ratios of the standard gasses, which should all be shown on this figure. Furthermore, Fig 6 and 7 show serious deviations of calibrated vs. true values for both <sup>13</sup>C and <sup>18</sup>O between circa 2000 – 10000 ppm, of as much as 2 permil, even though those differences disappear at higher values. This deviation is unacceptable given the requirements of analyzing soil CO<sub>2</sub>, where differences of 2 permil may be highly significant from an ecological perspective. I note that CO<sub>2</sub> concentrations < 10,000 ppm are commonplace in most soil profiles, especially in shallow horizons that typically dominate CO<sub>2</sub> production, such that capacity for accurate and precise measurements in this lower concentration range is really critical. Even greater variability is shown in Fig 7, which appears to reach 4 per mil. This is not acceptable for natural abundance work.

Response: As per the suggestions made by both of the referees, we have conducted another round of calibration by diluting CO<sub>2</sub> gas using synthetic air instead of N<sub>2</sub>. Diluting the CO<sub>2</sub> standard gas with N<sub>2</sub> resulted in a standard deviation of 8.1(‰ for  $\delta^{13}\text{C}$  values and 4.7 (‰ for  $\delta^{18}\text{O}$  values respectively. Diluting CO<sub>2</sub> standard gases with synthetic air resulted in a standard deviation of 6.44(‰ and 6.818(‰ for  $\delta^{13}\text{C}$  and  $\delta^{18}\text{O}$  respectively (see Fig.1a-b). With our new calibration curves (see Fig.1 c-d, & Table.1,2), we are able to bring down standard deviation to 0.08(‰ for  $\delta^{13}\text{C}$  and 0.04(‰ for  $\delta^{18}\text{O}$  (see Fig.2a-b (residual distribution), Fig.3a-d (Corrected  $\delta^{13}\text{C}$  and  $\delta^{18}\text{O}$  values). By introducing the new calibration correction (see Fig.3) the values are very close to the target value across the whole concentration range and hence we are

C2

confident that the system is suitable for ecosystem studies based on measuring subtle changes in isotopic signature of CO<sub>2</sub> across plant soil atmosphere continuum.

2. Also, basic details about the soils investigated are missing that are necessary to interpret the measured values of  $\delta^{13}\text{C}$  and  $\delta^{18}\text{O}$  of CO<sub>2</sub>. For example, what are the carbonate concentrations and isotope ratios in the calcareous soil, and how do they vary with depth? What are the  $\delta^{13}\text{C}$  values of SOM? This is a prerequisite for interpreting the soil profile CO<sub>2</sub> values. Also, to calculate the isotope ratios of soil-respired CO<sub>2</sub>, we need measurements of the atmospheric boundary condition. See Davidson 1995 GCA, doi:10.1016/0016-7037(95)00143-3. Note that several recent papers neglected have reported  $\delta^{13}\text{C}$  of CO<sub>2</sub> from soil profiles using high temporal-resolution optical measurements, these should be discussed or at least mentioned.

Response: We have measured soil samples for bulk  $\delta^{13}\text{C}$ , bicarbonate  $\delta^{13}\text{C}$  &  $\delta^{18}\text{O}$  values and also % of carbon content in the soil across a depth profile of (0-80 cm) for the calcareous soil (See Fig.4a-c and Fig.5). We observed an increase in  $\delta^{13}\text{C}$  values (of bulk soil and carbonate) in deeper soil layers (See Fig.4 a,c). This fits to our assumption of an increased contribution of bicarbonate derived  $\delta^{13}\text{C}$  enriched CO<sub>2</sub> in deeper soil layers. Our aim is to establish a method which enables continuous online measurement of soil gas  $\delta^{13}\text{C}$  &  $\delta^{18}\text{O}$  values at different soil depths and hence calculating the isotope ratios of soil-respired CO<sub>2</sub> is not done in this manuscript. This would be beyond the scope of a calibration focused paper – we however show the importance to properly calibrate laser based systems to obtain valid measurements of  $\delta^{13}\text{C}$  and  $\delta^{18}\text{O}$  of soil CO<sub>2</sub> which is a prerequisite for assessing the rate and isotopic composition of soil respiration.

3. I am skeptical as to the validity of the temperature tests employed. Note that we need to know the temperature of the analyte gas itself, which may be substantially different than the temperature of the water bath through which it circulates unless the residence time of the gas in the tubing and the heat transfer properties of the tubing allow for sufficiently rapid temperature equilibration, which may not completely occur if flow rates

C3

are high. For example, certain applications require heating of gasses at a sampling inlet to avoid condensation, yet the temperature of the gas at the point of the analysis may be substantially different (e.g. -4 – 40C) for some other optical gas analyzers, and should optimally be controlled within the analyzer cavity itself. Thus, unless the exact temperature of the gas at the point of measurement can be determined, I would not trust the results from the water bath experiment. Regardless, details of the analysis flow rate should be reported (and whether these rates were controlled during sample analyses – TMC's are mentioned for standards only).

Response: The laser spectrometer was able to maintain the temperature inside the optical cavity quite stable at 46.61°C irrespective of the fluctuations in the gas temperature (See Figure attached below). It is clear that the temperature maintained in the water bath will not get directly reflected in the sample gas due to multiple reasons including diffusion barrier of the PTFE tubing and higher flow rates, never the less, there will be an increment or decrement in the gas temperature. The aim is to show that the system is also stable in field conditions where temperature fluctuation is happening. The system is running in a closed loop meaning there is enough time for the gas for heat exchange. We have adjusted the part where test for equipment stability under fluctuating temperature is done in the modified manuscript.

4. There are numerous issues with grammar, style, and errant capitalization throughout. The figures and tables have a strange mix of fonts (be consistent!) and the legends are compressed. Please follow standard procedures for presenting your MS (provide captions as text in the document, not as images). There is a significant typo in Table 1.

Response: This is addressed and rectified in the revised manuscript. We will also let a native speaker do the final editing of the manuscript.

5. Finally, it should be noted that the useful temporal resolution of the measurements will never actually be 1hz as reported given the Allan variance results.

Response: Not exactly clear what the reviewer meant by "it should be noted that the

C4

useful temporal resolution of the measurements will never actually be 1hz as reported given the Allan variance results". It is always useful to get a better temporal resolution which can be used for identifying short term dynamics of CO<sub>2</sub> efflux (e.g., diurnal pattern of soil CO<sub>2</sub> efflux). Meaning more data points are available for taking an average across a time frame which is best for noise correction by using Allan deviation curves.

6. Was water vapor removed from the analyte gas, and if so, how? Response: Yes, water vapor was removed using drierite desiccant cartridges. We will add this information to the revised manuscript.

Figure.1: Deviation of measured  $\delta^{13}\text{C}$  and  $\delta^{18}\text{O}$  of CO<sub>2</sub> (a,b) when diluted using synthetic air. (c-d) shows diff-  $\delta^{13}\text{C}$ , diff- $\delta^{18}\text{O}$  values across a concentration gradient. Red and Blue dots shows measured  $\delta^{13}\text{C}$  &  $\delta^{18}\text{O}$  values of two different gases with distinct isotopic signatures, red and blue dashed lines represents absolute  $\delta^{13}\text{C}$  &  $\delta^{18}\text{O}$  values of the respective gases. Black line denotes model fit for diff-  $\delta^{13}\text{C}$ , diff- $\delta^{18}\text{O}$  values across changing CO<sub>2</sub> concentration (300 – 25000 ppm).

Figure 2: Residual distribution of modeled data for diff-  $\delta^{13}\text{C}$ , diff- $\delta^{18}\text{O}$  values across changing CO<sub>2</sub> concentration (300 – 25000 ppm).

Figure 3: Corrected  $\delta^{13}\text{C}$  &  $\delta^{18}\text{O}$  values of two different standard gases measured after correcting for concentration dependent drift.

Figure.4: Bulk  $\delta^{13}\text{C}$  (a), bicarbonate  $\delta^{13}\text{C}$  (c) and % of carbon (b) in soil across a depth profile of (0-80 cm).

Figure.5: bicarbonate  $\delta^{18}\text{O}$  in soil across a depth profile of (0-80 cm).

Please also note the supplement to this comment:

<https://www.soil-discuss.net/soil-2018-9/soil-2018-9-AC1-supplement.pdf>

Interactive comment on SOIL Discuss., <https://doi.org/10.5194/soil-2018-9>, 2018.

C5

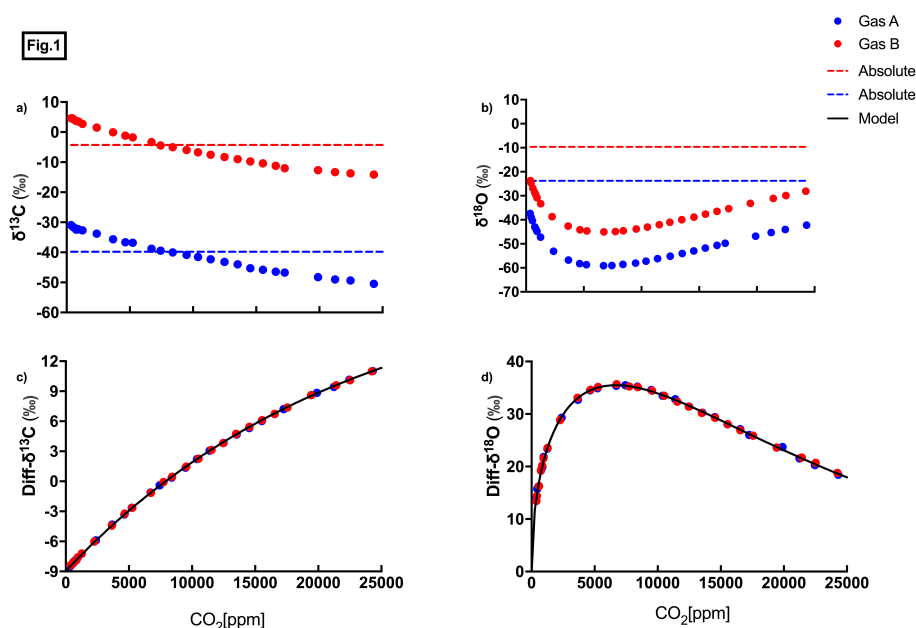
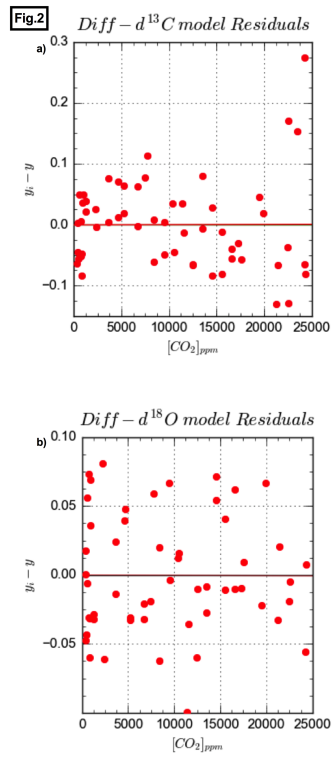


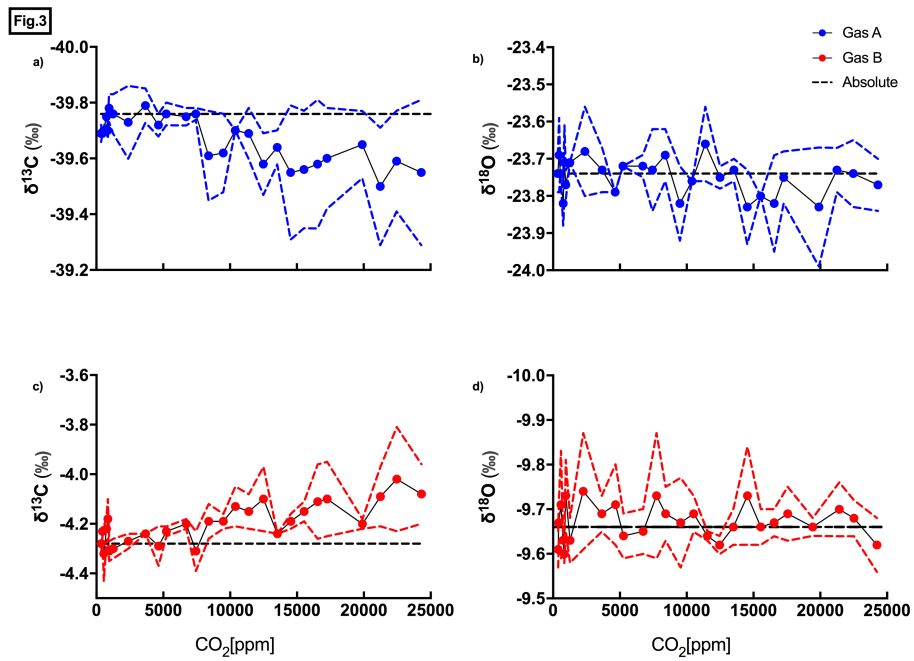
Fig. 1.

C6



**Fig. 2.**

C7



**Fig. 3.**

C8

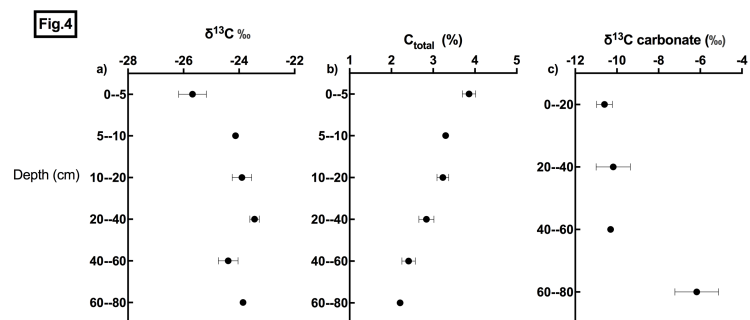


Fig. 4.

C9

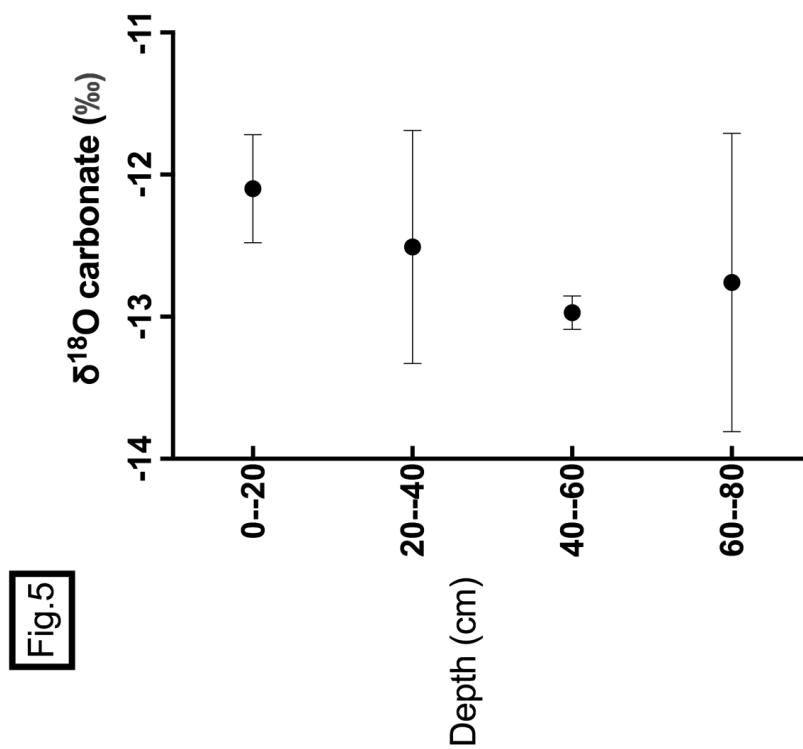


Fig. 5.

C10

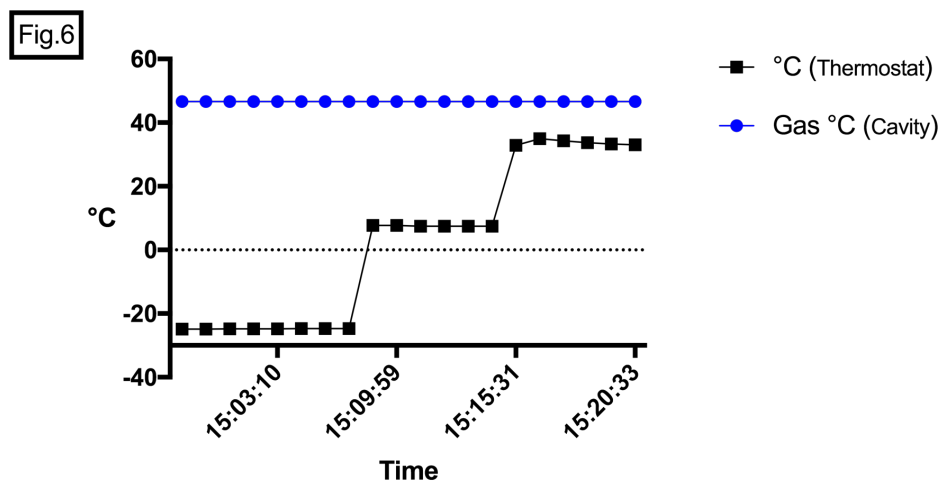


Fig. 6.